Properties of Hydrates Data Sheet

Student Name:

Partner Name:

Date	Instructor's Initials	Grade

Part One

Mass of Crucible = 8.923 g

Part Two

Name of Compound = Cobalt(II) Sulfate

Initial Color	Color after Heating	Solution Color

Initial Mass of Sample + Crucible	Final Mass of Sample + Crucible	Final Sample Mass
14.684 g	12.179 g	

Calculate the %Mass Lost. Show your work.

Observations

Name of Compound = Cobalt(II) Chloride		
Initial Color	Color after Heating	Solution Color
Initial Mass of Sample + Crucible	Final Mass of Sample + Crucible	Final Sample Mass
13.625 g	11.797 g	
Calculate the %Mass Lost. Show your work.		
Observations		
Name of Compound =		
Initial Color	Color after Heating	Solution Color
Initial Mass of Sample + Crucible	Final Mass of Sample + Crucible	Final Sample Mass
9.897 g	9.893 g	
Calculate the %Mass Lost. Show your work.		
Observations		

Part Three

Mass of Crucible = 8.763 g

Name of Compound = Copper(II) Sulfate

Initial Color	Color after Heating	Solution Color

Initial Mass of Sample + Crucible	Final Mass of Sample + Crucible	Final Sample Mass
13.687 g	11.915 g	

Calculate the %Mass Lost. Show your work.

Observations

Moles of Anhydrous Compound	Moles of Water Lost	Ratio of Water Moles to Compound Moles

Show your calculation of the moles of anhydrous ionic compound.

Show your calculation of the moles of water.
Show you calculation of the ratio between water moles and compound moles. Write the complete molecular formula for the hydrate.
Comment on any sources of error in your technique How would they effect your results. How would you prevent them in the future?